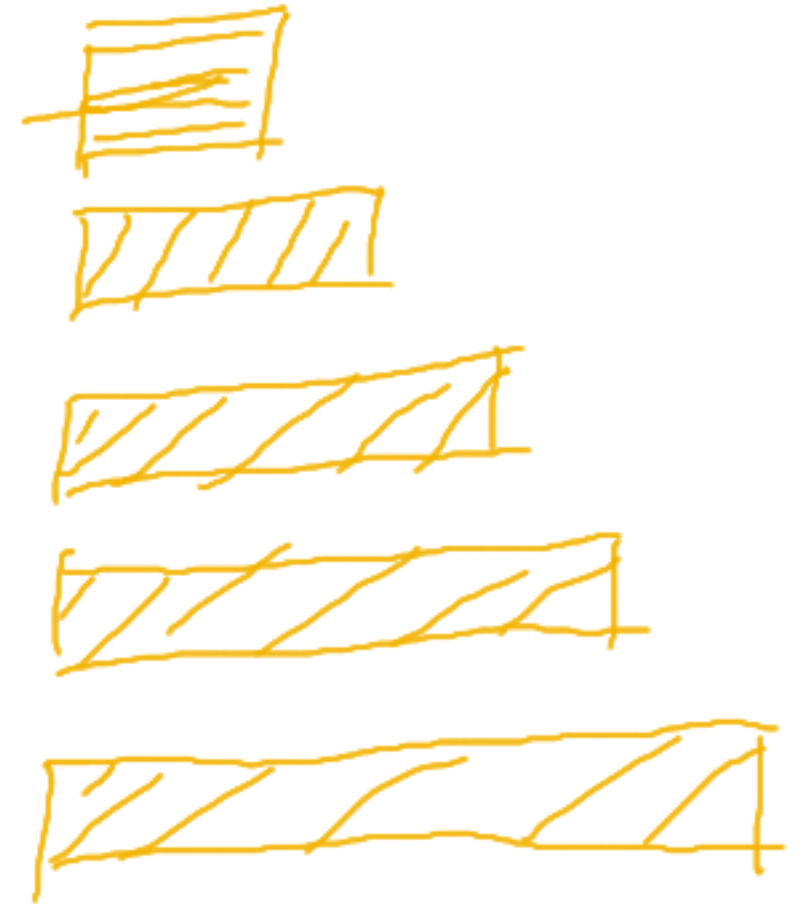


ইপডাৰ মাঙ



✓ চিত্ৰাৰ পিছত বাক্য

✓ স্মিট কামি চিত্ৰ কৰাটো আশা

✓ উত্তৰনী চিত্ৰ কৰা পাব

অন্যদে (যাক
আনন্দ হৈ

১০
২১
৩২
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✓ স্মৃতি কঠোর থাকবে।

✓ আনন্দ/সুখ দুই থাকবে।

✓ সূক্ষ্ম/ভিত্তিক কমাতে।

✓ আনন্দের মানসিক আরও উপকার হবে।

Hints for the problems of Chapter 38

Problem #1

Hints

Photon energy $E = hf$
 $= hc/\lambda$ — 34

$0.6 \text{ eV} = 6.63 \times 10^{-34} \cdot f$

light velocity $c = \lambda f$
 $\lambda = ?$

Problem # 2

$$\begin{aligned}\lambda &= 590 \text{ nm} \\ &= 590 \times 10^{-9} \text{ m}\end{aligned}$$

Kinetic energy

$$\frac{1}{2}mv^2 = hf$$

nm	10^{-9}	nanometer
μm	10^{-6}	micrometer
mm	10^{-3}	millimeter
m	\rightarrow	
<u>km</u>	10^{+3}	
<hr/>		
pm	10^{-12}	pico-meter
fm	10^{-15}	femto - "

$$\frac{1}{2}mc^2 =$$

$$\frac{hc}{\lambda}$$

$$v =$$

$$\sqrt{\frac{2hc}{\lambda m}}$$

m/sec.

#5

$$E = hf$$

$$= \frac{hc}{\lambda}$$

$$\lambda =$$

$$\frac{1}{1650763.73} \text{ m}$$

Photoelectric effect:

$$h\nu = \frac{KE_{\max}}{\frac{1}{2}mv_0^2}$$

+ ϕ ✓
work function
energy needed to dislodge an electron from an atom

$$\frac{dq}{dt} = I$$

$$I = \frac{Q}{R}$$

$$KE = z$$

$$E = \phi$$

$$z$$

$$hc \rightarrow 2.2 \text{ eV}$$

$$\lambda$$

$$eV_s =$$

$$\text{Kinetic Energy} = eV_s$$

$$\frac{\text{Energy}}{\text{charge}} = \text{potential}$$

#17

$$E = K_{\max} + \phi$$

$$\phi = 4.5 \text{ eV}$$

$$K_{\max} = E - \phi$$

$$E = 5.8 \text{ eV}$$

$$= 1.3 \text{ eV}$$

$$v = ?$$

$$\frac{1}{2}mv^2 = 1.3 \text{ eV}$$

$$v = ?$$

#27

Change in wavelength in Compton scattering

$$\Delta\lambda = \frac{h}{mc} (1 - \cos\phi)$$

$$\lambda' - \lambda = \frac{h}{mc} (1 - \cos\phi)$$

$$\phi = 30^\circ$$

$$\lambda = 2.4 \text{ pm}$$

$$\lambda' = ?$$

λ = incident wavelength of γ -ray
 λ' = outgoing " " the γ -ray

$$\lambda' > \lambda$$

c = velocity of light

m = electron mass

Subatomic particles

Radius $r \sim 10^{-12} \text{ m} \text{ --- } 10^{+5} \text{ m}$

picometer to femtometer range.

Example: [✓]proton, [✓]neutron, [✓]electron, [✓]atom

Matter wavelength
or de Broglie wavelength

$$\lambda = \frac{h}{p} = \frac{h}{mv}$$

Planck's constant
momentum

#46

Kinetic energy $\Rightarrow \frac{1}{2} m_e v^2 = 1000 \text{ eV}$

Momentum $p = \frac{m_e v}{2m_e}$

$p^2 = 2m \times 1000 \text{ eV}$

$p = \sqrt{2m_e \times 1000 \text{ eV}}$

Mass of electron

$m_e = 9.1 \times 10^{-31} \text{ kg}$

$\lambda = \frac{h}{p}$

p

#47

$$V = 25 \text{ keV}$$

Steps:

① Get KE

↓
② p (momentum)

↓
③ λ (matter wave)

We know

$$\frac{KE}{e} = V$$

$$KE = e \cdot 25000 \text{ eV}$$

$$= 1.6 \times 10^{-19} \times 25 \text{ K}$$